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Vodka and other high-proof alcoholic beverages typically don't freeze in a home freezer. Yet, the freezing point is higher than that of pure ethanol (-173.5°F or -114.1°C). Vodka may be considered a solution of ethanol (solute) in water (solvent). When considering freezing point depression, look at the freezing point of the solvent.

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Freezing Point Depression **Boiling Point Elevation and Freezing Point Depression Problems - Equation / Formula** ~~Molar Mass By Freezing Point Depression - Molality \u0026 Van't Hoff Factor - Chemistry Problems~~ Freezing Point Depression Experiment

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The equation for the freezing point depression of water is: 1 cc spoon 30-mL and 150-mL beakers Cup, clear plastic, 9 oz Digital scale Goggles Test tube 13 x 100 mm Thermometer Materials Not Included Δ77" Kym (Equation 6.1) where dTf is the freezing point depression; Kyis the freezing point depression constant for the solvent; and m is the molality (moles of solute per kilogram of solvent) of the solution.

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The freezing point depression (ΔT f), or the degrees by which the freezing point of the solution is lower than the freezing point of the pure solvent, can be calculated by Eqn. 1. ΔT f = K f * m Eqn. 1. Here, K f is the molal freezing point depression constant that is characteristic of the solvent. For

Freezing Point of Pure Water = 0°C Freezing Point of 20%NaCl/Water = -16°C The IMF of water are disrupted by the salt, reducing the energy required to separate the water molecules The decrease in freezing point of any solvent is based only on the amount of solute and the identity of the solvent *The identity of solute as no effect

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The freezing point depression can be predicted using the equation T=K f x m x i, where T is the change in freezing point, i is the number of ions in the solution per mole of dissolved NaCl (i = 2), m is the molality of the solution, and Kf is the molal freezing point constant for water which is 1.86oC/m.

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Lab 3 - Freezing Point Depression

LAB 6 Questions for Freezing Point Depression Table 6.1 Soluti on Lowest Temperature °C 1 st Trial °C Freezing Point 2 nd Trial °C Freezing Point Average °C Freezing Point ΔT f °C Sucro se-7.9-1.6-1.6 H 2 O-8.0 +0.0 +0.0 1. Was there supercooling? Would you expect the water or the sugar solution to have the most supercooling (see the ...

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The freezing point depression of Solution 1 is -4.4 C, and the freezing point depression of Solution 2 is -7.7 C. The freezing point depression can be determined by subtracting the temperature of the solution by the temperature of the water. The molality of Solution 1 is 2.4 mol / kg, and the molality of Solution 2 is 4.1 mol / kg.

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Plugging the m c into equation 2, you can calculate the freezing point depression, Δ T f, of the solution. Last, you must subtract this from the freezing point of the solvent to get the predicted freezing point of the solution. In Part A of this experiment, you will determine the freezing point of pure stearic acid.

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